Acids and Bases

There are three commonly used definitions for acids and bases –

* Arrhenius acids and bases
	+ Acids release hydrogen in aqueous solution
		- Example – HCl, HNO3
	+ Bases release hydroxide ion in solution
		- Example – NaOH
* Bronsted-Lowry Acids and Bases
	+ Acids are proton donors
		- Example – CH3OH
	+ Bases are proton acceptors
		- Example – NH3
	+ Bronsted Lowry definitions are more general than Arrhenius definitions
* Lewis Acids and Bases
	+ Acids are electron acceptors
		- Example: AlCl3
	+ Bases are electron donors
		- Example: C6H6
	+ Lewis definitions are the most general

Lewis Acids and Bases

Bronsted-Lowry Acids and Bases

Arrhenius acids and bases