Acids and Bases

There are three commonly used definitions for acids and bases –

* Arrhenius acids and bases
  + Acids release hydrogen in aqueous solution
    - Example – HCl, HNO3
  + Bases release hydroxide ion in solution
    - Example – NaOH
* Bronsted-Lowry Acids and Bases
  + Acids are proton donors
    - Example – CH3OH
  + Bases are proton acceptors
    - Example – NH3
  + Bronsted Lowry definitions are more general than Arrhenius definitions
* Lewis Acids and Bases
  + Acids are electron acceptors
    - Example: AlCl3
  + Bases are electron donors
    - Example: C6H6
  + Lewis definitions are the most general

Lewis Acids and Bases

Bronsted-Lowry Acids and Bases

Arrhenius acids and bases